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solution mol solute.  $1 \text{ L} = 1000$   
mL. The molarity of a solution is a  
ratio of the moles of solute per  
liters of solution. The units for  
molarity are written as mol/L or  
M. This measurement is used to  
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Ciinc. Solution  $M_1 \times V_1 = \text{mol}$  W.ter

Concenaiaed solution is diluiej

with more solvent. I him . snliiin.

$M_1 \times V_1 = \text{mol}$  USEFUL EQUATIONS

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$M_1 \times V_1 = M_2 \times V_2$  molarity =  
molsolutc L.solution 1 L- 1000 ml.  
A solution can be made less  
concentrated in a process called  
dilution. This is accomplished by  
adding more solvent.

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1 L = 1000 mL The molarity of a solution is a ratio of

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many grams of silver chromate  
will precipitate when 150. mL of

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0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?  $2 \text{ AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{ KNO}_3(\text{aq})$

0.150 L  $\text{AgNO}_3$  0.500 moles  $\text{AgNO}_3$  1 moles  $\text{Ag}_2\text{CrO}_4$

331 ...

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USEFUL EQUATIONS. molarity =  $\frac{\text{mol solute}}{\text{L solution}}$ . 1 L = 1000

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As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions

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occur in aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

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first. 1] How many grams of calcium phosphate can be produced from the reaction of 2.50 L of 0.250 M Calcium chloride with an excess of phosphoric acid?

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volume is then used to calculate  
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the sample using the  
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