

## Chapter 19 Electrochemistry Answers

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Chapter 19: Electrochemistry. oxidation. reduction. electric current. electrochemical charge. the loss of electrons, which also corresponds to an increase in... the gain of electrons, which also corresponds to a decrease in... The flow of electric charge. A device in which a chemical reaction either produces or is ca...

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Answer: E cell = -0.22 V; the reaction will not occur spontaneously. Applying the Nernst equation to a simple electrochemical cell such as the Zn/Cu cell discussed in Section 19.2 allows us to see how the cell voltage varies as the reaction progresses and the concentrations of the dissolved ions change.

Chapter 19.4: Electrochemical Cells and Thermodynamics ...  
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Chapter 19: Electrochemistry Dr. Pahlavan / Dr. Ghanbaripour 3 9. Given the following data:  $2\text{Fe}^{3+} + 2\text{e}^{-} \rightarrow 2\text{Fe}^{2+}$   $E^\circ_{\text{red}} = -0.44 \text{ V}$ ,  $\text{Ag}^{+} + \text{e}^{-} \rightarrow \text{Ag}$   $E^\circ_{\text{red}} = 0.80 \text{ V}$  Answer the following questions with respect to the reaction;  $\text{Fe}^{2+}(\text{aq}) + 2 \text{Ag}(\text{s}) \rightarrow \text{Fe}(\text{s}) + 2 \text{Ag}^{+}(\text{aq})$

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The hydrogen-oxygen fuel cell is described in Section 19.4 (a) What volume of  $\text{H}_2(\text{g})$ , stored at  $25^\circ\text{C}$  and  $1.00 \text{ atm}$ , would be needed to run an electric motor drawing a current of  $8.5 \text{ A}$  for  $3.0 \text{ h}$ ? (b) What volume (in liters) of air at  $25^\circ\text{C}$  and  $1.00 \text{ atm}$  will have to pass into the cell per minute to run the motor?

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Chapter 19 Electrochemistry Chang 4 19.4 SPONTANEITY OF REDOX REACTIONS Electrical work is needed to move a charge through a conductor: Electrical energy = charge x potential difference Units: Joules = Coulombs x Volts F is Faraday's constant; the electrical charge contained in 1 mol of e-is equal to 96500 C.  $1 \text{ F} = \text{mole}^- \text{ Coulombs } 1 \text{ 96500} = \text{V mole}^-$

CHAPTER 19 ELECTROCHEMISTRY  
Read Chapter 18: Entropy, Free Energy, and Equilibrium & Read Chapter 19: Electrochemistry Answer the following problems in the space provided. For problems involving an equation, carry out the following steps: 1. Write the equation. 2. Substitute numbers and units. 3. Show the final answer with units. There is no credit without showing work.

AP Chemistry Chapter 18 & 19: Thermodynamics ...  
Chapter 19 Electrochemistry Math Summary Relating Standard Cell Potential to Standard Half Cell Potentials  $E^\circ_{\text{cell}} = E^\circ_{\text{oxidation}} + E^\circ_{\text{reduction}}$  (standard conditions assume 1.0 M concentrations) Relating Half Cell Potentials when Written in Opposite Directions  $E^\circ_{\text{ox}} = -E^\circ_{\text{red}}$  for half reactions written in opposite directions

Chapter 19 Electrochemistry Math Summary  
Answer:  $\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^- \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}(\text{l})$ ;  $\text{Sn}^{2+}(\text{aq}) \rightarrow \text{Sn}^{4+}(\text{aq}) + 2\text{e}^-$  - The Pt electrode in the permanganate solution is the cathode; the one in the tin solution is the anode.